QB365 Question Bank Software Study Materials

Periodic Classification of Elements 50 Important 1 Marks Questions With Answers (Book Back and Creative)

10th Standard

Science

Total Marks : 50

Multiple Choice Question

50 x 1 = 50

1) The number of periods and groups in the periodic table are_____.

(a) 6,16 (b) 7,17 (c) 8,18 **(d) 7,18**

2) The basis of modern periodic law is_____.

(a) atomic number (b) atomic mass (c) isotopic mass (d) number of neutrons

3) _____ group contains the member of halogen family.

(a) 17^{th} (b) 15^{th} (c) 18^{th} (d) 16^{th}

4) _____ is a relative periodic property.

(a) atomic radii (b) ionic radii (c) electron affinity **(d) electronegativity**

5) Chemical formula of rust is _____.

(a) $FeO.xH_2O$ (b) $FeO_4.xH_2O$ **(c) $Fe_2O_3.xH_2O$** (d) FeO

6) In the alumino thermic process the role of Al is _____.

(a) oxidizing agent **(b) reducing agent** (c) hydrogenating agent (d) sulphurising agent

7) The process of coating the surface of metal with a thin layer of zinc is called_____.

(a) painting (b) thinning **(c) galvanization** (d) electroplating

8) Which of the following inert gases have 2 electrons in the outermost shell.

(a) He (b) Ne (c) Ar (d) Kr

9) Neon shows zero electron affinity due to _____.

(a) stable arrangement of neutrons **(b) stable configuration of electrons** (c) reduced size (d) increased density

10) _____ is an important metal to form amalgam.

(a) Ag **(b) Hg** (c) Mg (d) Al

11) The green layer found on the copper vessel is due to the formation of _____

(a) basic copper carbonate (b) cupric oxide (c) cuprous oxide (d) copper chloride

12) The number of neutrons in $_8O^{16}$ is _____

(a) 8 (b) 16 (c) 32 (d) 24

13) Modern periodic law is based on _____

(a) atomic mass **(b) atomic number** (c) number of neutrons (d) Both (a) and (b)

14) As the positive charge increases, the size of the cation _____

(a) decreases (b) increases (c) remains constant (d) First increases and then decreases

15) Electronegativity values are based on _____

(a) bond energy (b) electron affinity (c) ionisation energy **(d) all the above**

16) Electronegativity values of Na and CI are 0.9 and 3.0 respectively predict the nature of bonding.

(a) Ionic (b) Covalent (c) Coordinate (d) Metallic

17) The process of extracting the ore from the earth's crust is _____

(a) Metallurgy **(b) Mining** (c) Smelting (d) Leaching

18) Metals are _____

(a) Electro positive (b) Electronegative (c) both (a) and (b) (d) neither (a) nor (b)

19) Which among the following are the ores of aluminium?

(i) Bauxite

(ii) Cryolite

(iii) Corundum.

(a) Both (i) and (ii) (b) Only (c) Only (iii) **(d) (i), (ii) and (iii)**

20) The chemical formula of sodium meta aluminate is _____

(a) $NaAlO_2$ (b) Na_2AlO_2 (c) $NaAl_2O_2$ (d) $Na_2Al_2O_3$

21) The chief ore of copper is _____

(a) Copper pyrites (b) Copper glance (c) Cyprite (d) Rupy copper

22) Blister copper contains _____

(a) 50% pure copper (b) 99% pure copper and 1% impurities **(c) 98% pure copper and 2% impurities**

(d) 75% pure copper and 25% impurities

23) First period contains _____ elements.

(a) 4 **(b) 2** (c) 5 (d) 3

24) Fourth and fifth period called as_____ period.

(a) short (b) shortest **(c) longer** (d) longest

25) First period is the _____ period.

(a) short **(b) shortest** (c) long (d) longest

26) Atomic number from 3 to 10 are called as _____.

(a) first **(b) second** (c) third (d) fourth

27) Atomic number from 19 to 36 are called as _____.

(a) first (b) second (c) third **(d) fifth**

28) Atomic number from 55 to 86 are called as _____.

(a) first (b) second (c) fifth **(d) sixth**

29) The lanthanides and actinides which form part of Group 3 are called_____ elements

(a) S block (b) P block (c) Transition elements **(d) Inner transition elements**

30) Group 18 called as _____.

(a) Alkali metals (b) Alkaline earth metals (c) Halogens **(d) Noble gases**

31) Group 16 called as _____.

(a) Alkali metals (b) Alkaline earth metals **(c) Oxygen (or) Chalcogen family** (d) Rare gases

32) When a neutral atom gain an electron, it becomes a negatively charged ion called _____.

(a) Cation **(b) Anion** (c) Neutral ion (d) all the above

33) Oxide ores are purified by this method.

(a) Gravity separation method (b) Magnetic separation method (c) Froth floatation method

(d) Chemical method (or) leaching

34) Lighter ores such as sulphide ores are concentrated by the following method.

(a) Froth floatation method (b) Magnetic separation method (c) Gravity separation method (d) Chemical method

35) Matte is a mixture of _____.

(a) $Cu_2S + FeS$ (b) CuS + FeS (c) $Cu_2O + FeS$ (d) CuO + FeS

36) The concentrated ore is roasted in _____ of air.

(a) Roasted in excess of air (b) Roasted in less air (c) Roasted in absence of air (d) None

37) The formula for malachite green is _____.

(a) $CuCO_3.Cu(OH)_2$ (b) $CuCO_3.Al(OH)_3$ (c) $CuCO_3.Fe_2O_3$ (d) $CuCO_3.Fe(OH)_3$

38) _____is the second most abundant element in the earth crust.

(a) Aluminium **(b) Iron** (c) Copper (d) None

39) The chief ore of Iron is _____.

(a) Aluminium hydroxide **(b) Ferric oxide** (c) Ferrous oxide (d) Ferrous sulphide

40) The chemical formula for the rust is _____.

(a) $Fe_2O_3.xH_2O$ (b) $Fe_2O_3.H_2O$ (c) $Fe_3O_4.H_2O$ (d) $FeO.H_2O$

41) Iron with 2 - 4.5% of carbon is called _____.

(a) Pig iron (b) Steel (c) Wrought iron (d) None

42) Statues are made up of _____ Alloy.

(a) Cu, Sn (b) Cu, Zn (c) Al, mg, Mn (d) Al, mg

43) Stainless steel is an alloy of _____.

(a) Aluminium **(b) Iron** (c) Copper (d) None

44) Iron alloys are also called as _____.

(a) Ferrous alloys (b) Ferrous alloys (non) (c) Ferric alloys (d) Non Ferric alloys

45) The rocky impurity associated with an ore is called _____

(a) mining **(b) matrix** (c) flux (d) slag

46) The_____ is a less reactive metals.

(a) mercury (b) sodium (c) Aluminum (d) Calcium

47) _____ is used in making aeroplanes and other industrial machine parts.

(a) copper (b) Iron (c) Silver **(d) Aluminium**

48) When copper reacts with dil HNO_3 _____ gas is liberated.

(a) Nitric oxide (b) sulphur oxide (c) copper oxide (d) carbon di oxide

49) When fusing the metal by melting of zinc and copper _____ alloy is collected.

(a) brass (b) bronze (c) magnalium (d) duralumin

50) When a sodium atom loses an electron it forms Na^+ ion. The radius of Na^+ ion is lesser than Na atom. This is because,

(a) The attractive force of nucleus is more in Na^+ ion than Na atom

(b) The attractive force of nucleus is more in Na atom than Na^+ ion

(c) Number of protons present in Na atom is less than Na^+ ion

(d) Number.of electrons present in Na^+ ion is more than Na atom